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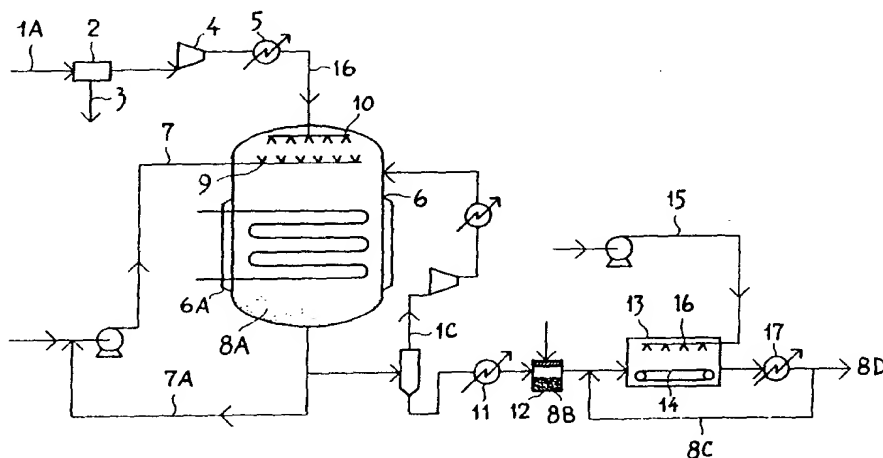
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(54) Title: METHOD FOR PRODUCTION OF GAS HYDRATES FOR TRANSPORTATION AND STORAGE

**(57) Abstract**

Method for the production of gas hydrates for storage and transportation of hydrate forming gases. The gas is pressurized and cooled, whereupon the gas (1b) is supplied to a vessel (6), to which water (7) is added simultaneously to the vessel (6) to form fine water droplets that are dispersed in the gas (1b) supplied. The temperature and pressure in the vessel (6) are adjusted to produce hydrate from water and gas. The gas (1b) supplied and the water droplets (7) react to form hydrate (8a); then the hydrate is then withdrawn from the reactor (6), optionally cooled in a cooling apparatus (11), and then passed to an agglomeration step (12) to agglomerate the hydrate (8a), thus increasing the density of the hydrate (8a) and to embed more gas in the interstices between the hydrate particles (8a). The agglomerated hydrate particles (8a) are transported to a suitable transportation unit or storage container to transport/store the hydrate at adiabatic conditions at atmosphere pressure or at a slight gauge pressure at a temperature below 0 °C, preferably at minus 10 °C to minus 15 °C.

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Method for production of gas hydrates for transportation and storage.

The present invention concerns a method as stated in the introductory of patent claim 1, for the production of gas hydrates stable for storage, particularly hydrates of natural gas or associated natural gas, for onshore and offshore transportation or for the storage of the same.

Background.

In cases where a pipe system for transporting gas, either natural gas or gas mixed with other hydrocarbons is unavailable, it is difficult to carry out an efficient use. In such cases, continuous combustion of the gas is impossible, it cannot be used on the spot, e.g. at an offshore platform, or transported to the customer through a pipeline.

One possibility in such situations is to re-inject the gas to the reservoir to promote the oil production. The development of certain isolated gas fields, e.g. offshore, is economically impossible without permitting utilization on the spot or transportation through a pipeline.

It is also possible to carry out treatment on the spot by producing liquidous natural gas, methanol and ammonia. However, these three possibilities require extensive treatment of the natural gas and/or another gaseous product and requires equipment that is comprehensive and on a scale that is not suited for offshore production.

US Patent 3,514,274 discloses a method for solving the transportation problem, in which natural gas is converted to hydrates and transported/stored in propane or other C₄-C₅ hydrocarbons. In this case, the propane is used as a recyclable energy carrier, and the natural gas hydrate is dehydrated at the delivery point and converted to pure natural gas simultaneously with converting the propane to propane hydrate. Then, the propane hydrate can be used again for the production of natural gas hydrate, in which compressed and cooled natural gas is contacted with propane hydrate in a reactor, thus converting propane hydrate to propane carrier liquid and natural gas to natural gas hydrate. However, this method has the disadvantage that dead weight, i.e. propane, must be transported all the time. Moreover, the transportation and storage temperature must be as low as -22°C to avoid evaporation of the propane carrier liquid.

NO 149976, patent publication laid open to public inspection, discloses a method for transporting natural gas in a boat submerged in water. Natural gas and fresh water is separately fed into a submerged marine ship and combined to form gas hydrate, whereupon the hydrate is kept stable during transportation by means of the

5 hydrostatic pressure and the relatively low temperature of the sea water. However, this method requires that the pressure and the temperature is maintained throughout transportation, and requires use of specially constructed submarines that have a low loading capacity compared with surface ships. Moreover, problems may arise if the shipping terminals are located in an area with shallow and relatively warm water.

10 Moreover, the method disclosed in this patent publication is naturally limited to marine transportation only, and will not contribute to a solution where only pipeline transportation or transportation as LNG is available.

In other cases, one faces gas emission such as polluting by-products from industrial processes where a removal of the gas is desired but purification or

15 destrucion on the spot is impossible.

Object

The main object of the present invention is to provide a method for the treatment of hydrate forming gases, such as natural gas or natural gas mixed with or enclosed

20 in other hydrocarbons or water, or polluting gases or gas to be supplied to an industrial or biotechnical process that permits economically satisfactory storage, transportation and use of the gas without using pipeline or immediate transportation by tankers or tank cars, and without the need for use of pressure or any carrier liquid during transportation or storage.

25 Another object of the invention is to provide a method that in addition is environmentally acceptable and that can be realized with an acceptable risk with respect to security and economy.

Brief description of the invention

30 The principle of the invention is stated in the characterizing part of patent claim 1. Further beneficial features of the invention are stated in the remaining patent claims.

The present invention concerns a method for the production of storage stable gas

hydrates from water and hydrate forming gases, such as CO_2 , H_2S , natural gas and associated natural gas, just to mention a few. However, in the following, natural gas is in general described as the gaseous component in the production process, but it should be evident that a person skilled in the art can apply the principle of the

5 invention to consider hydrate forming gases other than natural gas, and the invention should for that reason not be regarded as limited to use of natural gas only. The present method for production of gas hydrates can be adapted to both onshore and offshore operation.

At the production site, oil and water is separated from the natural gas and natural
10 gas mixed with other hydrocarbons, whereupon the purified gas is compressed and cooled. The condensed gas produced by this compression and cooling is removed in a separator, in which temperature and pressure is adjusted to produce predetermined hydrocarbons, preferably butane and higher hydrocarbons. The separated cooled gas is further compressed and passed through a heat exchanger and cooled.

15 The compressed gas is then guided to a reactor vessel and, together with pressurized water, expanded to a lower pressure through nozzles or the like, thus producing fine droplets dispersed in the expanded natural gas. The water and the gas will react almost immediately to produce natural gas hydrate comprising frozen water with enclosed gas. The pressure and temperature conditions in the reactor are
20 adjusted to favor hydrate formation, and the gas pressure prior to expansion is preferably adjusted to provide cooling during expansion by means of the Joule-Thomson effect. Provided that there is a defined ratio between pressure and temperature that represents equilibrium between gas hydrate and water, the reactor temperature is preferably decreased a few degrees below the equilibrium
25 temperature, thus increasing the reaction rate for the formation of natural gas hydrate. A sub-cooling from 1 to 10°C is in most cases sufficient, and a typical sub-cooling varies from 2 to 6°C .

The natural gas hydrate, formed as fine powder, is transported out of the reactor either by the reactor gauge pressure or by means of a mechanical transportation
30 apparatus. Any excess gas is then separated from the hydrate powder, e.g. in a cyclone, whereupon the separated gas is compressed, cooled and recirculated back to the hydrate reactor. The hydrate powder is then cooled partly by ordinary heat loss

during flow in the transportation pipe and partly through expansion to a lower pressure and optionally further cooling in a heat exchanger. The cooled hydrate powder is then optionally transferred to an agglomeration step, such as pressing or pelletizing, to provide a more dense natural gas hydrate and to embed further gas in pores. The resulting hydrate particles can then optionally be provided with a protective ice shell by spraying them with water, whereupon the water will freeze and form ice. In cases where the hydrate particles are unable to provide proper cooling to the water droplets added, further cooling must be provided, e.g. by cooled gas flowing through the wetted hydrate particles. The ice shell will provide more fracture strength and thermal insulation. If desired, the ice shell can also be strengthened with reinforcing materials, such as fibers, to further strengthen the ice shell and therefore the hydrate particles.

The hydrate particles are then cooled to a suitable storage temperature, and the particles can be stored or transported stable for a longer time, up to several weeks, at adiabatic conditions and at a pressure near atmospheric pressure. At a later stage when the gas is to be used, heat is supplied to the natural gas hydrate to decompose same to form gas and ice. The water can, if desired, be recycled or discharged, without any environmental risks. However, it is an advantage to recycle the water back to the gas hydrate production process, firstly because the water itself represents a low temperature reservoir, and secondly because the water, provided that it is kept at a temperature below $+30^{\circ}\text{C}$, still contains seeds that promotes the reaction rate for the hydrate formation, as further described below.

Transportation and storage

The gaseous hydrate particles can be used for storage and transportation of gases. They can also be used for operating transporting means onshore and offshore. Other gases may also be used to produce the gaseous hydrate particles. These other gases can be commercial products or pollutants or other gas types that form in natural or industrial processes. Gaseous hydrate particles can be used in power stations and in processes intended for reduction of pollution. Gaseous hydrate particles can be used where gas has to be added in large amounts, in aquatic environments, both natural and artificial.

The gaseous particles can be stored in offshore platforms in subsea vessels under pressure. These vessels can be located on the sea bed or adjacent to the platform. They can be pressurized hydrostatically with a water column through a valve arrangement with manometer to keep the vessel and the sea water separated by means of a water column. The gaseous particles can be stored as solid material in gas or surrounded by cooled water or a hydrocarbon based liquid. In addition to subsea vessels, tankers, barges and the like can be used, or submerged vessels made up of a stiff or flexible material.

Hydrate particles with embedded gas can be transported from offshore storage vessels by boat, tankers, barges or floating containers towed by tugboats to the shore. In the most preferred arrangement, hydrate particles are pumped from the storage vessels offshore through a pipeline to a tanker. The tanker can, but does not need to, be able to store the particles under gauge pressure. The particles can be transported to the shore as solid cargo or in water or in a hydrocarbon based liquid. Gas that escapes from the particles during transportation can be pressurized and/or used to operate the tanker and the cooling equipment.

Hydrate particles can also be stored in underground storage rooms, such as large caverns blown in rock formations. This can be accomplished by cooling/refrigerating the underground storage cavern prior to the supply of gas hydrates, so that any naturally occurring water freezes and forms an isolating ice shell on the "vessel" walls. In this way, gas escape from the storage cavern can be prevented. Like ordinary isolated vessels, the gas hydrate produced in accordance with the invention can be stored near atmospheric pressure, as described in further detail below.

The hydrate particles with embedded gas are after the transportation pumped or transferred by other ways from the tanker to one or several storage tanks onshore. The particles melts and the gas can escape. The melting can be accomplished using different types of heating, e.g. with emission from a gas operated power station. Cold melting water can be used as coolant for any power station, thus making the ordinary cooling towers redundant.

When the tanker is emptied, melting water and process water can be loaded. The water can have its origin from a former cargo. The melting water will be ballast for the tanker from the shore to an offshore platform. When the tanker loads the

particles at the platform, the melting water is unloaded. The vessels at the platform accept the melting water for use in the hydrate production. If desired, air may be removed from the melting water and the process water and optionally pre-treated. The air removal can be effected onshore and/or offshore. In addition, the water can
5 be used for injection to a reservoir.

Transferring hydrate particles from e.g. the production unit to a storage vessel or transportation vessel can be accomplished by using pneumatic transportation systems. The carrier gas is in this case preferably cooled natural gas as opposed to air that is used in ordinary pneumatic transport systems. The use of cooled natural gas in such
10 systems will cool the hydrate particles during transportation and thus contribute in a positive manner to particle stability.

Detailed description of the invention

In the following, the invention is described in further detail with reference to the
15 drawings, in which

Figure 1 is a hydrate equilibrium diagram for a typical treated natural gas for use with the present invention,

Figure 2 is a simple diagram illustrating a general method for production of hydrates in accordance with the invention,

20 Figure 3 is a simplified process flow sheet that illustrates the method for production of hydrate powder in accordance with the present invention,

Figure 4 illustrates an alternative method for providing a protecting ice shell on hydrate particles,

Figure 5 illustrates schematically an experimental arrangement to measure the
25 storage stability of gas hydrates, and

Figure 6 and 7 show temperature change and amount of gas emitted as a function of time during the testing of storage stability of natural gas hydrate in accordance with the experimental arrangement in Figure 5.

30 Figure 1 shows a pressure/temperature diagram for a typical treated natural gas, applied as an example in the method in accordance with the present invention, the diagram provided with a equilibrium curve for hydrate. The gas in the example

- comprises, after the removal of heavier hydrocarbons, 92% methane, 5% ethane and the remainder propane. However, the treated gas can nevertheless contain small amounts of other gases, such as carbondioxide, oxygen or air, without adversely effecting the subsequent production of hydrates. As appears from the curve, it is not
- 5 necessary that the formation temperature for hydrate be lower than 0°C. According to Figure 1, the formation pressure for natural gas hydrate is 104 bar at +20°C, whereas the formation pressure at 0°C will be about 8 bar. Hydrate formation will occur at the high pressure side/low temperature side of this curve. Water can establish two different lattice types, the first having an empiric formula of 8X •
- 10 46H₂O (where X represents one gas molecule) and the second 24X • 136 H₂O, provided that all cavities in the crystal lattice are occupied by gas. These crystals form a lattice having relatively large cavities that are occupied by gas. Accordingly, the gas is not directly bonded to the water molecules in the crystal, but is merely captured by the crystal's geometric limitations until the lattice structure breaks down.
- 15 The heat of formation for gas hydrate is exothermic and, to keep the temperature in the reactor system constant, the reaction heat must be removed from same, partly by means of gas expansion cooling and partly by means of indirect or direct cooling of the reactor.

In theory, gas hydrates are unstable at atmospheric pressure, and even at -15°C a

20 pressure of e.g. at least 4.5 bar is required to keep the hydrate exemplified in Figure 1 in a stable state. To disintegrate a hydrate into its respective components, it is required to supply the hydrate with its dissociation heat, and will accordingly assume a meta stable state at adiabatic conditions in a cooled state, even at pressures close to atmospheric pressure. Experiments carried out in connection with the present

25 invention have demonstrated that natural gas hydrate is stable even at an ambient temperature of -1.5°C. The bulk mass of hydrate will serve as isolation itself and only hydrate particles located close to the vessel walls will recieve heat from the surroundings. In this way, bulk cargo of hydrate can be transported/stored in a stable state, provided that the vessel is thermally insulated with respect to the surroundings

30 and that the temperature is kept at a low level. This meta stable state can be further improved by compressing the hydrate particles and optionally providing them with an external protecting ice shell.

Figure 2 illustrates in general a method for production of storage stable gas hydrates in accordance with the present invention. In process step 1, gas is pre-treated, e.g. by removing heavier hydrocarbons from natural gas, and thereafter in process step 2 is supplied to a reactor 2 together with water pre-treated in process step 1. The gas and the water react in accordance with the equilibrium conditions in question for hydrate formation and forms gas hydrate, in most cases with a snow-like appearance. The formed gas hydrate is then transported to process step 2, in which any unreacted gas and water is removed from the formed hydrate particles, whereupon the hydrate particles optionally are compressed/agglomerated and provided with a protecting ice shell. The formed and optionally post-treated hydrate particles are then further transferred in process step 5 to a transportation or storage container, in which storage or transportation occurs at conditions close to adiabatic and at a pressure close to atmospheric. The hydrate can then be stored for a longer period of time or transported for long distances without the risk of the hydrate decomposes into its respective components.

In the following, a general method for production is presented with a discussion of the different aspects of this method with reference to Figure 3.

Natural gas and natural gas mixed with other hydrocarbons is separated from oil and water (not shown) at the production site. The purified gas is transferred through a compressor and cooled in a heat exchanger cooled by air or sea water. The condensed gas 1 produced from this compression and cooling step is removed in a separator 2; temperature and pressure is adjusted to produce predetermined hydrocarbons 3, preferably butane and higher hydrocarbons. These isolated liquidous components 3 can be used in combustion processes and in operation of platforms. The separated, cooled gas is compressed in a compressor 4 and passed through a heat exchanger 5 and cooled, e.g. by air or sea water.

The compressed and cooled gas is transported to a reactor vessel 6 that is also supplied with pressurized water 7, to produce gas hydrate 8a having embedded gas. This pressurized water 7 is supplied to the reactor through nozzles 9 or the like and expanded to a lower pressure and to a temperature that results in formation of gas hydrate 8a. The water pressure is not critical for the formation of gas hydrate, and the pressure can be adjusted to a desired level provided that the pressure is

higher than the reactor pressure. However, the water pressure should be adjusted to achieve sufficient volumetric injection of water to the reactor and properly dispersing the water in the gas phase as fine droplets. Because of the exothermic character of the formation reaction, it is preferred that the construction and the operation

5 conditions, particularly the pressure, is chosen to provide best possible cooling of the feed streams. This is accomplished by adjusting the pressure of the gas 1b supplied to the reactor vessel 6 to provide cooling by expansion (Joule-Thomson effect), and the expansion is in this case also carried out by means of nozzles 10. Moreover, it is favourable to adjust the reactor temperature a few degrees below the hydrate

10 equilibrium temperature, generally 1-10°C, preferably 2-6°C, thus increasing the rate of hydrate formation. The formation rate can also be increased by adding small seeds of hydrate crystals to the water to be supplied to the reactor so that hydrate can more easily grow from these in the reactor. The formation of hydrate nucleus occur at the interface between the water and the gas bulk phase. The water 7 is preferably

15 dispersed as thoroughly as possible in the gas bulk phase. The water can be supplied to the reactor 6 through the same openings, e.g. nozzles, as the gas, thus establishing a mixing effect at the supply location in the reactor. Moreover, water droplets in the reactor 6 can be dispersed in the gas bulk phase by, for example, means of a spreader means, such as a rotating plate with nozzles distributing fine droplets

20 (preferably having a diameter of the order micrometer), or by using physical guiding or blocking means inside the reactor, or by using a stirrer (not shown). Recirculated unreacted gas can also be supplied to the reactor perpendicular to the main flow of fresh gas feed, thus achieving even better mixing of the reactants. However, the reactor pressure and the respective initial pressures for gas and water can be

25 determined as desired, depending on the total pressure loss in the system and the gas pressure available. With respect to the process heat balance, a general rule says that the lower the reactor pressure, the less energy is required to produce gas hydrates based on the total energy content in the hydrate. On the other hand, the reaction rate for the formation of gas hydrate will increase with the pressure, and accordingly the

30 reactor pressure must also be adjusted in view of the type of gas supplied to the reactor.

The gas hydrate formed in solid state (snow like appearance) is then transferred

out from the reactor vessel, e.g. by means of a mechanical transportation apparatus or by means of the reactor gauge pressure. The hydrate particles 8a are separated from any unreacted gas, and liquidous water is removed. The pressure downstream of the reactor is controlled optionally by adjusting the operation pressure in the pipe
5 connecting the reactor and the separator. For example, at an operation pressure of 50 bar, a pressure downstream the reactor of 20 bar would be suitable. Excess gas is preferably separated from the formed hydrate particles in one or more cyclones, or similar apparatus for the separation of solid matter from a fluidum, whereupon the hydrate optionally is transferred to a suitable apparatus 12 for agglomerating the
10 particles, e.g. by drum treatment and stirring, pressing, extruding, heat treatment and drying, or liquid suspension, of which drum treatment, pressing and extruding are the preferred methods, as described in further detail below.

Before the water is supplied to the reactor, it can be ventilated to remove oxygen and other gases (not shown). The water can be treated with stabilizing agents,
15 additives and/or supplied with small seeds of hydrate crystals (as stated above). The stabilizing agents increase the storage and transportation ability of the hydrate particles with embedded gas. These agents may be produced on the spot from hydrocarbon fractions separated from the starting material, either from natural gas or natural gas together with other hydrocarbons. The additives can be compounds that
20 decrease the surface tension of water, thus increasing the reaction rate for the formation of gas hydrate.

As set forth above, the hydrate forming reaction is exothermic, but the contribution from expansion of gas by utilizing the Joule-Thomson effect to the total cooling requirement is small. For example, the heat of formation for natural gas
25 hydrate from the composition stated in Figure 1 at temperatures above ca. 0°C is about 2075 kJ/kg. Accordingly, the hydrate reactor must be cooled, either directly or indirectly. Direct cooling can for example be provided by circulating excess gas through an external refrigeration plant. In such cases, a need for an additional compressor will arise. Indirect cooling 6A can be accomplished with a cooling jacket
30 or cooling elements, e.g. provided with a coolant from a closed circuit cooling system in the form of a refrigeration unit.

The mass and energy balance of the stream supplied to the reactor vessel is

preferably adjusted to convert the substantial part of the water to hydrate particles, thus operating the process with excess gas. The reactor vessel can also be operated with excess water, and then, water must be separated away. The process can also have gas and water in excess. However, operating the reactor with excess gas is
5 preferred. In this way, dry hydrates are formed that will decrease the risk of accumulation of hydrate and blocking of the reactor outlet.

Minor amounts of gas and any water can flow along with the hydrate particles. The unreacted and removed components of gas and flowing water can be recirculated; water 7a is recirculated and combined with the fresh water feed 7, and
10 separated gas 1c is compressed, cooled and passed directly back to the reactor. Compressing the recirculated gas to a pressure a little above the reactor pressure is sufficient so that the gas easily flow into the same. The recirculated streams can also be treated with additives and further treated with respect to production of hydrate particles (not shown). Unreacted gas from the reactor is optionally compressed and
15 supplied to another similar system operated at a higher pressure.

The hydrate particles with emedded gas are transported, as described above, optionally to equipment agglomerating or collecting the small particles to larger particles. The first hydrate particles are cooled and/or refrigerated in a refrigeration unit 11 prior to entering the agglomeration step 12. Cooling and freezing can be
20 accomplished by pressure change, direct supply of cooled/refrigerated gas and/or indirect heat exchange. The purpose of the agglomeration is to agglomerate the hydrate to decrease its volume and simultaneously provide volume for gas storage in the particle pore volume. The compression or "agglomeration" can occur at pressure and temperature conditions chosen to achieve an optimum gas content and particle
25 stability, i.e. the pressure and temperature must be at the high pressure side/low temperature side of the equilibrium curve for hydrate formation (Fig. 1). Additives can be mixed with the hydrate particles to improve their properties. Depending on the process conditions chosen, the total mass percent of gas can in general be in the range from 10 to 40 percent of the particle weight. After the agglomeration, the
30 hydrate particles 8b can be cooled and/or refrigerated, thus keeping the total gas content inside the particle. The diameter of the compressed hydrate particles varies with the method used for agglomeration and the degree of compression desired, but a

typical particle diameter for agglomerated natural gas hydrate particles is for example 2-20 mm. Likewise, the density will vary with the agglomeration method and degree of agglomeration, but a typical density is e.g. in the range from 850 to 950 kg/m³.

5 If desired, the agglomerated hydrate particles are transported to an apparatus 13 that covers the gas impregnated particles with a pure ice shell by spraying the particles with water that freezes and forms an ice shell on the particles. For example, this can be accomplished by spraying the agglomerated particles 8b with water 15 via nozzles 16 whereas the particles are transported downstream by means of a mass
10 transporter 14, e.g. a conveyor. The hydrate particles covered by ice are then cooled in a cooling apparatus 17. The ice shell thickness may be varied as required, but in general it is sufficient that the ice shell has a thickness from 0.5 to 1.5 mm. This process step of covering the hydrate particles with ice can be accomplished in several steps to further stabilize the hydrate particles by recirculating the partly ice-covered
15 hydrate particles in stream 8c back to the same operation 13, or transporting the same to a following step (not illustrated). Cooling in the cooling apparatus 17 can for example be accomplished with a cooled methane based mixture at a pressure and a temperature outside the conditions favouring the hydrate formation. The ice shell has two major effects on the stability of a hydrate particle. Firstly, diffusion of gas from
20 inside the particle to the environment is prevented because diffusion of gas through ice is negligible. Secondly, the ice shell provides a protecting shell that withstands a higher internal pressure from the particle. It can be verified that a spherical ice shell (pure ice) having a diameter of 15 mm and a shell thickness of 1 mm is able to withstand an internal pressure of about 5 bar. This pressure is in theory sufficient to
25 prevent a typical natural gas hydrate from decomposing at temperatures below -13°C at atmospheric pressure. However, experiments carried out in connection with the present invention has revealed that hydrates are stable even at temperatures as high as -1.5°C, but the stability will of course increase with decreasing temperature. To improve this effect further, the ice shell is optionally provided with reinforcing
30 materials, such as fibers. The ice strength increases with decreasing temperature and with the use of fiber reinforcement. The fiber material can also be supplied at the first particle production by addition to the pressurized and cooled water or in other

ways, e.g. by adding hydrate particles to the fiber material followed by mixing in a mixing unit, prior to the water spraying step. Moreover, the fiber material is optionally added in the agglomeration step when producing larger hydrate particles from the smallest gas filled hydrate particles. The produced, agglomerated and
5 cooled hydrate particles 8d, optionally provided with ice shell, are then ready for transportation or storage.

An alternative mode of providing an ice shell on the hydrate particles is illustrated in Figure 4. In this method, the hydrate particles 20 formed are wetted by spraying with water 21 in, for example, a separate chamber 22. The wetted hydrate particles
10 23 are then guided into the top of a tower 24 and brought in direct contact with cooled gas 25 downstream that is supplied to the tower 24 at the bottom of same. The cooled gas, e.g. natural gas, cools the wetted hydrate particles to effect freezing of the water to establish a protecting ice shell on same, whereupon the hydrate particles covered by ice are removed from the tower in stream 26. This method
15 provides for proper cooling of the wetted hydrate particles and results in a homogenous ice shell on the individual hydrate particles.

Such gas containing hydrate particles can be produced at offshore platforms or onshore. The platforms can be temporary or permanent. Onshore, the hydrate particles can be produced at a location close to hydrocarbon sources or other
20 locations. The gas supplied in this way can be natural gas or natural gas together with other constituents. It can also be pollution gas to be transported away for further treatment.

Example 1

25 This example illustrates one alternative method for the production of hydrate from natural gas by using the method of production in accordance with the invention, in which a relatively high reactor pressure of about 50 bar is applied.

Natural gas or associated gas is compressed and treated to remove components heavier than methane, ethane and propane in a manner known *per se*. The resulting
30 mixture comprises 92% methane, 5% ethane and 3% propane (mole percent).

The treated gas mixture having the composition set forth above is then compressed to about 100 bar, supplied to a hydrate reactor through a nozzle and expanded to a

pressure of about 50 bar. At the same time, water having a temperature of about 10°C is compressed to about 100 bar and supplied to the reactor by expansion through separate nozzles, thus forming small droplets that disperses in the expanded gas phase. However, the Joule-Thomson cooling from the expansion of the gas from 100 to 50 bar constitutes only 43 kJ/kg, i.e. about 2% of the total cooling requirement, and the remaining cooling requirement is realized by using external cooling; cooling jacket and cooling elements supplied with recirculated liquidous propane coolant including cooled compressed recycle gas. The temperature and the pressure at the reactor inlet is 13°C and about 50 bar, respectively, and according to the equilibrium curve for this composition (Figure 1), this condition is located just inside the hydrate forming area.

The natural gas hydrate formed, having a snow like consistence, falls down toward the reactor bottom by the force of gravity and exits the reactor to an environmental pressure of about 10 bar. The individual hydrate particles then have a density of about 920 kg/m³ and a gas content corresponding to 160-170 std. m³ pure natural gas per m³ hydrate powder, and comprise of about 15 mass percent natural gas and the remainder water. Typically, the particle size is from 1 to 10 mm. The hydrate powder is withdrawn from the reactor by the gauge pressure in the reactor, whereupon unreacted gas and water is separated from the gas hydrate formed, prepressurized, cooled and transferred back to the reactor 6; the volume stream of the recirculated gas is about 10 times as great as the amount of fresh gas fed to the reactor. The hydrate is then cooled to -15°C and compressed/agglomerated by pressing in a hydraulic press to a resulting particle size of about 5-15 mm, thus providing more embedded gas.

The produced agglomerated natural gas hydrate is then transported by means of cooled natural gas to storage vessels or to a transportation vessel. The cooled natural gas cools the natural gas hydrate by direct contact during transportation to a temperature of about -15°C, a temperature sufficiently low for this type of hydrate. The cooled natural gas hydrate is stored/transported in well insulated containers, preferably provided with a refrigeration unit, at a pressure close to atmospheric. The meta stable natural gas hydrate remains stable at these adiabatic storage conditions and withstands storage and transportation for several weeks without the need for

converting back to pure natural gas.

Example 2-4

These examples were provided to illustrate storage stability at different
5 temperatures for the gas hydrates in accordance with the present invention.

Example 2

Laboratory tests were carried out to test the thermal stability of a natural gas hydrate produced from the same natural gas composition as set forth in example 1.
10 This natural gas hydrate was produced in a testing batch reactor at high pressure and moderate temperature. The hydrate was removed from the reactor and cooled to minus 5°C in such manner that no decomposition occurred; i.e. the natural gas hydrate did not decompose during transportation from the reactor to the equipment used for measuring the thermal stability. The equipment is illustrated in Figure 5.
15 The solid hydrate 35 was located in a test tube 31 inside a closed container 32 and located in a container 34 at a constant temperature of minus 5°C. The closed cylinder 32 was kept at atmospheric pressure and connected with equipment for volumetric registration of any natural gas evolved/emitted from the gas hydrate. The test tube 31, the closed cylinder 32 and the external container 34 at constant
20 temperature were constructed to maintain almost adiabatic conditions in the test tube; i.e. heat was neither removed from nor added to the test tube. To the bottom of the test tube 31 a temperature gauge 36 was attached to measure the temperature in the gas hydrate.

The solid hydrate was stored in the test tube at minus 5°C for a long period of
25 time. The solid hydrate was stable and gave no indication of decomposing to gas and ice; i.e. no gas emission from the test tube was measured.

Example 3

The test tube 31 from example 2 above including the solid hydrate 35 and the
30 closed container 32 was moved to another container 34 having a constant temperature. This second container 34 had a temperature of +5°C. The closed container 32 and the test tube were gradually heated, and the solid hydrate started to

decompose to gas and liquid water. The results from these experiments are illustrated in Figure 6. The vertical axis to the left shows the percentage of gas emitted during the heating process. The last two hours of the storage period at near adiabatic conditions is illustrated in the figure; i.e. a storage temperature of minus 5°C and no
5 emission of natural gas. When the test tube and the surrounding cylinder were removed from minus 5°C to plus 5°, the temperature in the test tube started to increase, whereupon the natural gas hydrate started to decompose. The test tube temperature is shown in Figure 6 and is an approach to the real temperature in the solid natural gas hydrate, as appears from the construction of the testing apparatus in
10 Figure 5. All gas was removed within 6 hours (from 2 to 8 hours in Figure 6). The approached temperature in the test tube increased relatively fast (during about 0.5 hours) from minus 5°C to about 0°C. The temperature was constant at about 0°C, whereupon the natural gas hydrate decomposed (melted), as shown by the amount of natural gas evolved. When all natural gas was melted, the test tube temperature
15 increased to about +5°C, the same temperature as in the surrounding container 34.

Example 4

Another experiment was carried out by using the natural gas hydrate 35 produced and treated in the same manner as in Example 2, and the same storage temperature
20 was used (-5°C). After having stored the solid hydrate 35 at this temperature for a longer period of time without observing any emission of natural gas from the hydrate, the test tube 31 and the surrounding cylinder 32 were moved to a third container 34 having a constant temperature of about +20°C. The results of this experiment is shown in Figure 7. When the test tube 31 and the surrounding cylinder
25 32 were moved to a higher temperature (here indicated as 2 hours), the temperature in the test tube 31/35 started to increase, and the solid hydrate 35 started to decompose (melt). About 2 hours later (shown as 4 hours in Figure 7), all natural gas had escaped. During this period of time the test tube temperature increased rapidly to about +5°C. When the solid hydrate was melted, the test tube temperature
30 increased to about +20°C (not shown).

As appears from Figures 6 and 7, a natural gas hydrate can be heated at different rates, and the duration of the heating period affects the rate of decomposition; i.e.

the rate for gas recovery, e.g. when unloading hydrate from a tanker to a terminal onshore. The amount of gas recovered will be the same for different heating rates.

Other similar experiments have been carried out in connection with the present invention. They show that a typical natural gas hydrate can be stored in a stable state at a temperature below the normal water freezing point; a test with a surrounding temperature of minus 1.5°C resulted in stable hydrate at adiabatic conditions. A storage temperature of -5°C was used in Examples 2-4 above. Other storage temperatures (freezing temperatures) are also applicable. The storage temperature chosen will depend on the application in question; i.e. what extent of hydrate stability is required. It should be understood that the stability of the solid hydrate that is stored may be dependent on the storage temperature; a lower storage temperature results in a more stable gas hydrate. A gas hydrate that is to be stored for a longer period of time should be stored at a lower temperature than a hydrate that is to be stored for only a short period of time.

Claims.

1. Method for the production of gas hydrates for transportation and storage of gases, such as gases of the type natural gas or natural gas mixed with other hydrocarbons or water; pollution gas or gas to be supplied to an industrial or biotechnical process, the gas being pressurized and cooled, and combined with
5 pressurized water,

c h a r a c t e r i z e d in that

the gas (1b) is supplied to a vessel or a reactor (6), whereupon the water (7), optionally with added stabilizing agents or other additives, is simultaneously supplied to the vessel (6) in such manner that the water (7) is converted to small droplets that
10 disperses in the gas (1b) supplied, the temperature and pressure in the vessel (6) being adjusted to promote formation of hydrate from water and gas,

whereupon the gas (1b) supplied and the water droplets (7) react to form hydrate (8a) whereupon the resulting hydrate (8a) is withdrawn from the vessel (6) as any unreacted gas or water is separated from the hydrate particles and is recirculated
15 back to the vessel (6), and optionally cooled in a cooling apparatus (11), and further to an agglomeration step (12) for agglomerating the hydrate (8a) thus increasing its specific weight and to embed further gas in the interstices between the different hydrate particles (8a),

whereupon the agglomerated hydrate particles (8b) are transferred to a
20 transportation unit or storage container where the hydrate particles are transported/stored at atmospheric pressure or by a slight gauge pressure in cooled state, at near adiabatic conditions, at a temperature below 0°C, preferably at minus 10°C to minus 15°C.

2. Method in accordance with claim 1,
25 **c h a r a c t e r i z e d** in that the water (7) is supplied to the vessel (6) through at least one nozzle (9) or similar openings.

3. Method in accordance with claim 1 or 2,
c h a r a c t e r i z e d in that gas and water is supplied to the vessel through the same nozzles or openings.

4. Method in accordance with any of claims 1 to 3,
c h a r a c t e r i z e d in that the agglomerated hydrate particles (8a) produced are provided with a protective ice shell in at least one separate treatment step (13) to further improve the fracture strength of the hydrate particles by spraying the particles
5 with water that freezes to an ice shell on the hydrate particles (8a), whereupon the ice covered hydrate particles are cooled in a final cooling step (17).

5. Method in accordance with claim 4,
c h a r a c t e r i z e d in that the hydrate particles are provided with an ice shell by wetting them with water and bringing them in direct countercurrent contact with
10 upward rising cooled gas thus freezing the water supplied to a procecting ice shell.

6. Method in accordance with claim 5,
c h a r a c t e r i z e d in that the hydrate particles and/or the ice shell is strengthened with a reinforcing material, such as fibers, supplied to the hydrate particles via the water that wets same.

15 7. Method in accordance with any of claims 1 to 6,
c h a r a c t e r i z e d in that the water is treated with stabilizing agents, such as separated hydrocarbon fractions, or small hydrate crystal seeds, prior to the expansion of the water into the vessel.

8. Method in accordance with any of claims 1 to 7,
20 c h a r a c t e r i z e d in that pressure and temperature in the gas prior to the supply to the reactor (6) and pressure and temperature in the reactor (6) are adjusted to provide a cooling of the gas through the expansion into the reactor (6).

9. Method in accordance with claim 8,
c h a r a c t e r i z e d in that the reactor is cooled to a temperature in the range
25 from 1 to 10°C below the equilibrium temperature to produce gas hydrate at the pressure in question, preferably 2 to 6°C below the same.

10. Method in accordance with any of claims 1 to 9,
c h a r a c t e r i z e d in that the agglomeration of gas hydrate (8a) produced in the reactor (6) is carried out by drum treatment and agitation, pressing, extruding,
30 heat treatment and drying, or suspending in a liquid, preferably drum treatment, pressing and extruding.

11. Method in accordance with any of claims 1 to 10,
c h a r a c t e r i z e d in that the water being supplied to the reactor is provided with additives to decrease the water surface tension, thus increasing the reaction rate for the production of gas hydrate.

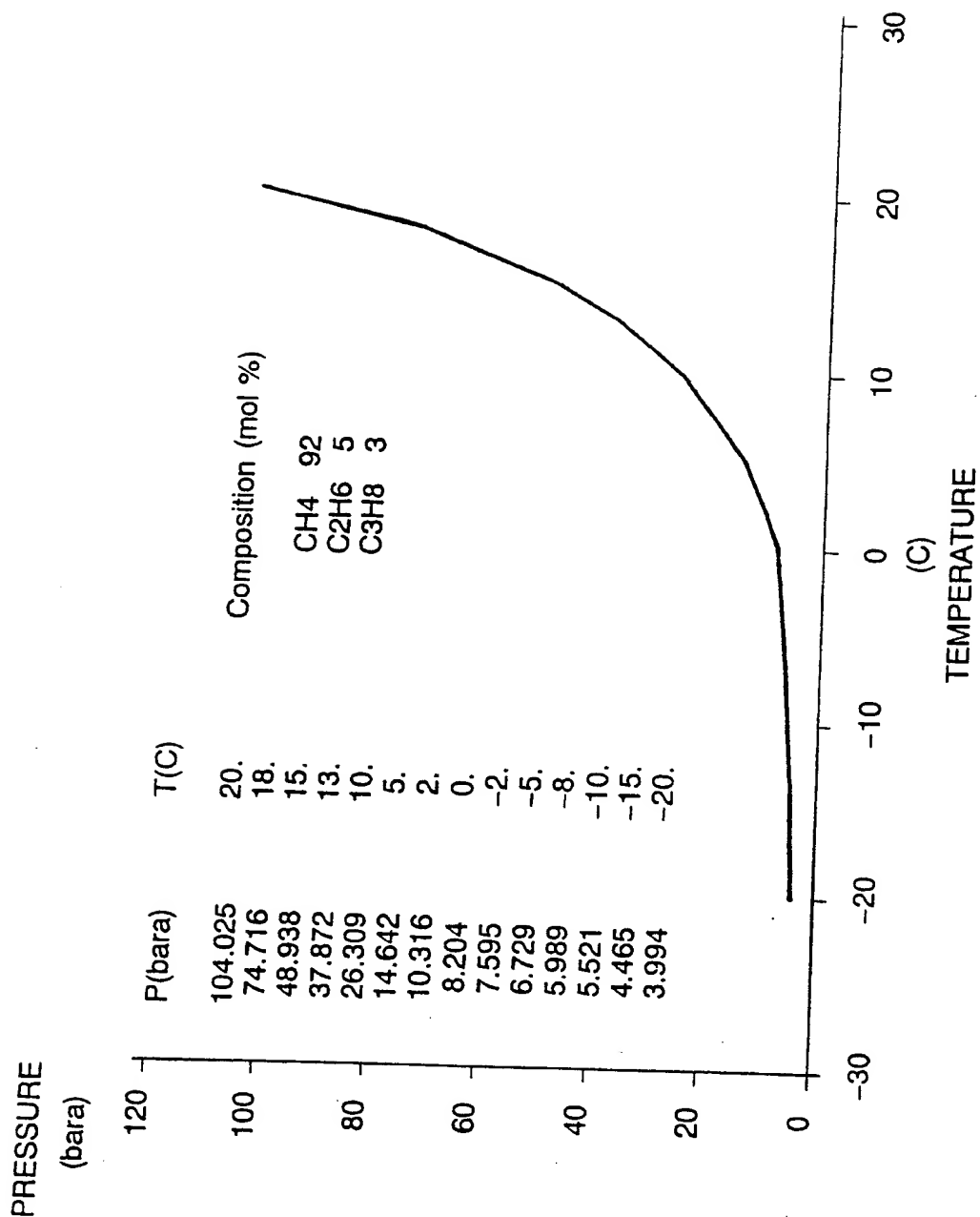
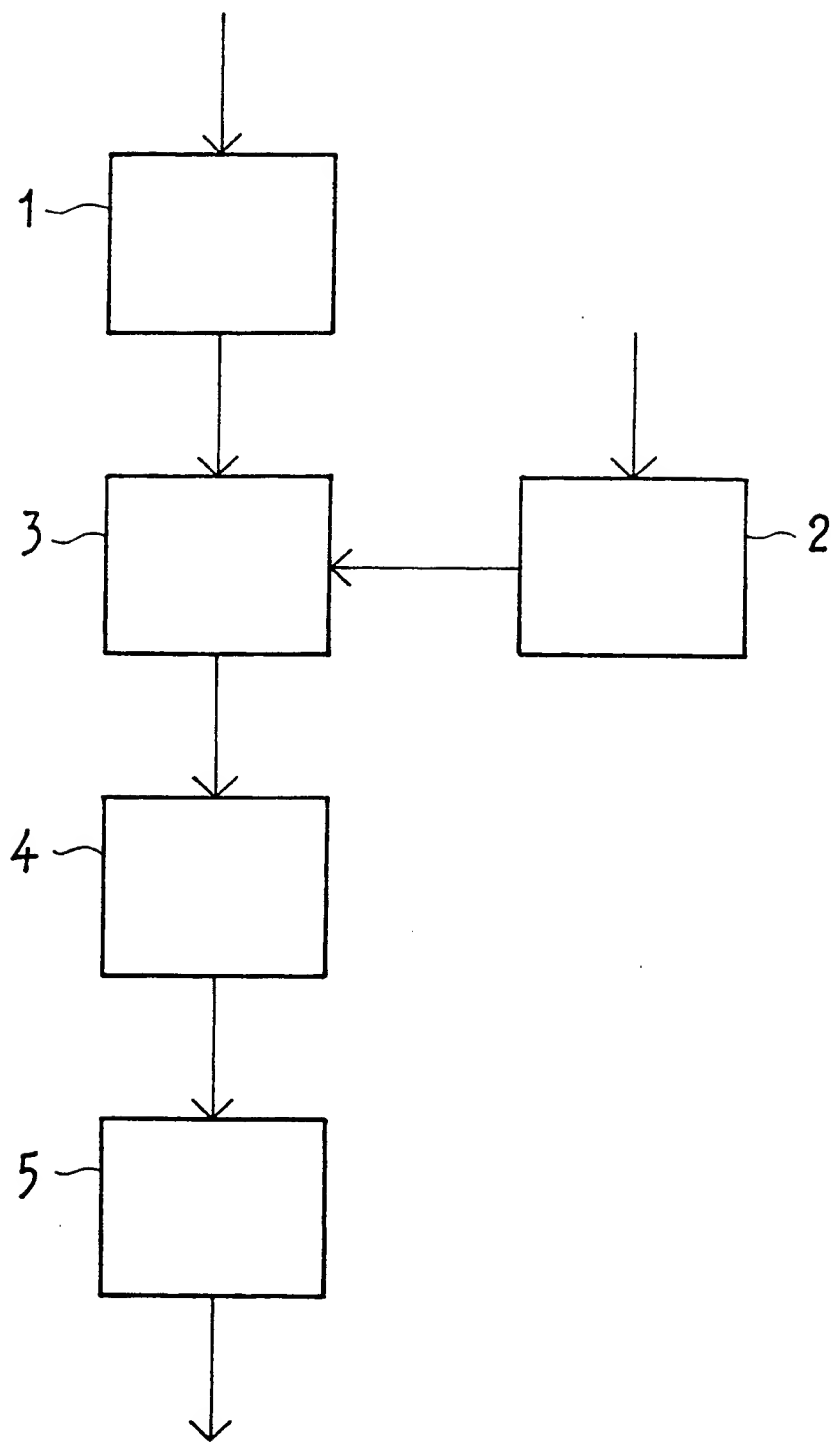


Fig. 1

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**Fig. 2**

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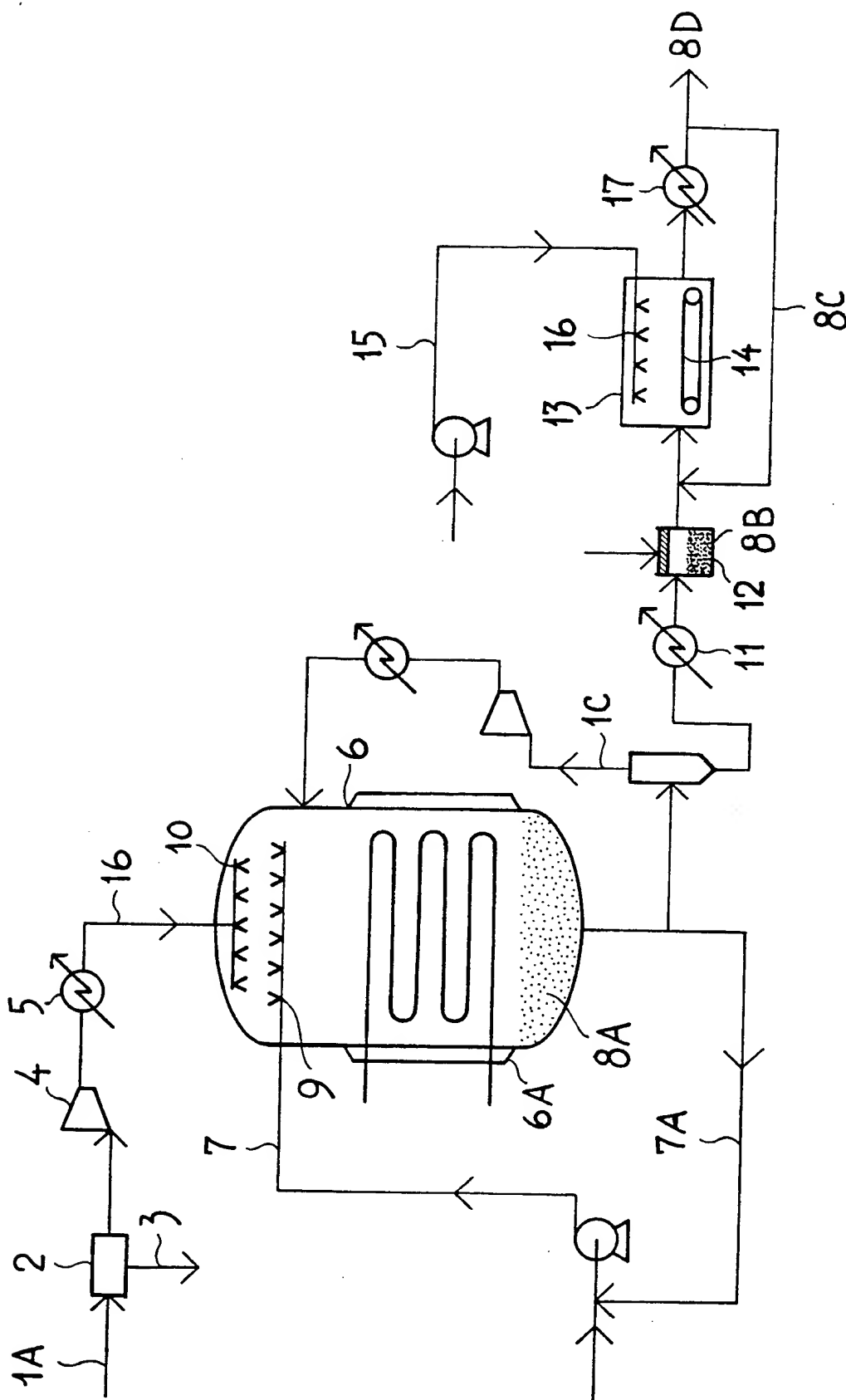
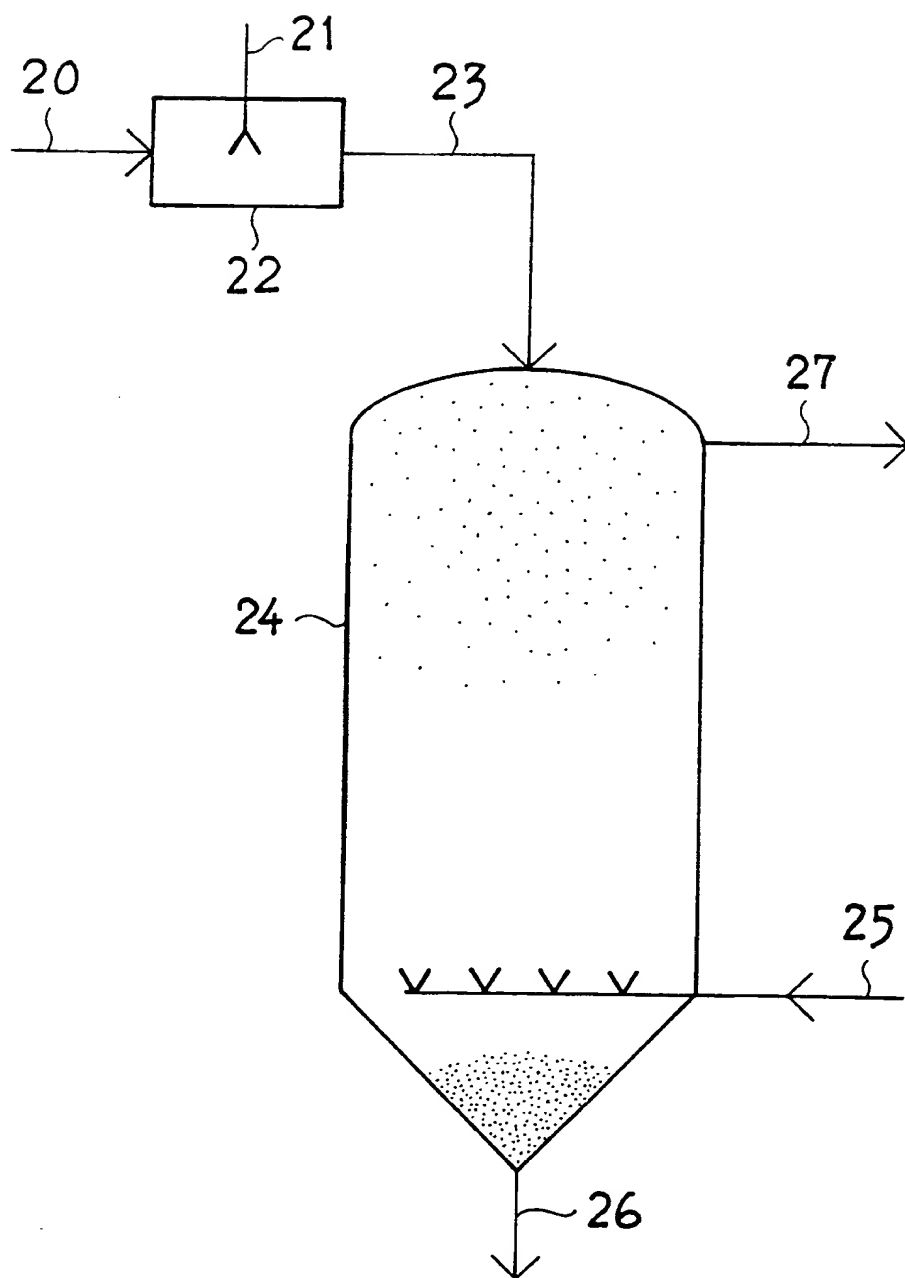
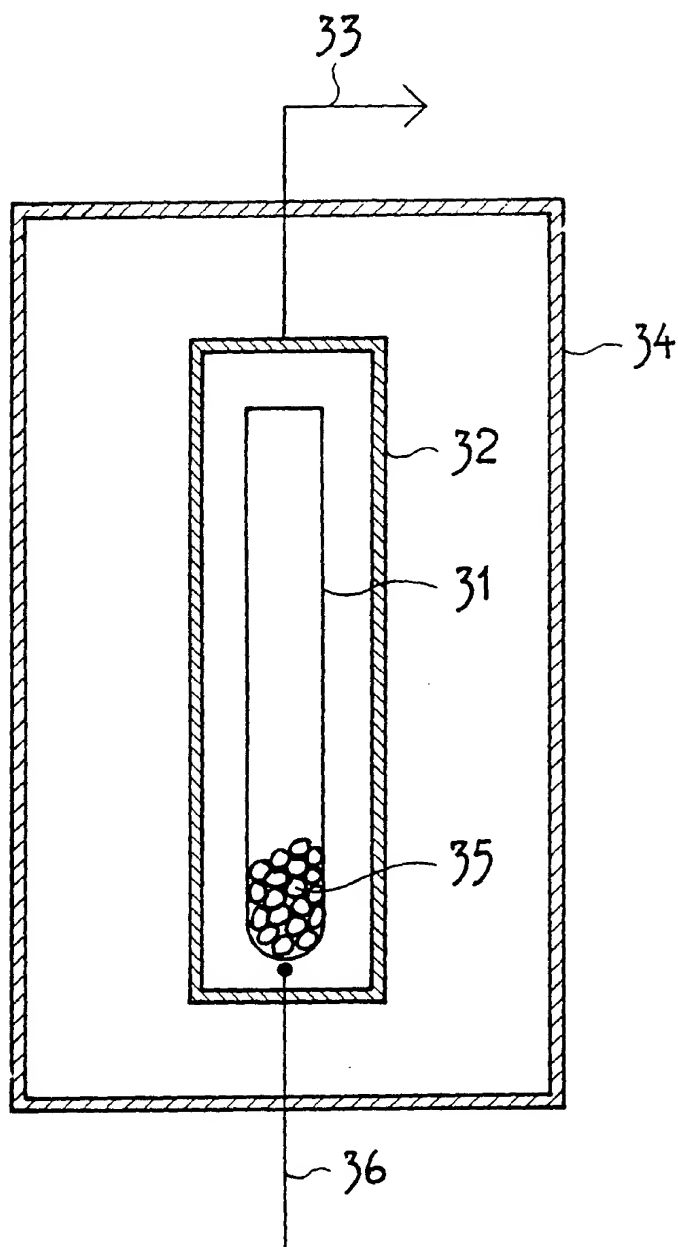


Fig. 3

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**Fig. 4**

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**Fig. 5**

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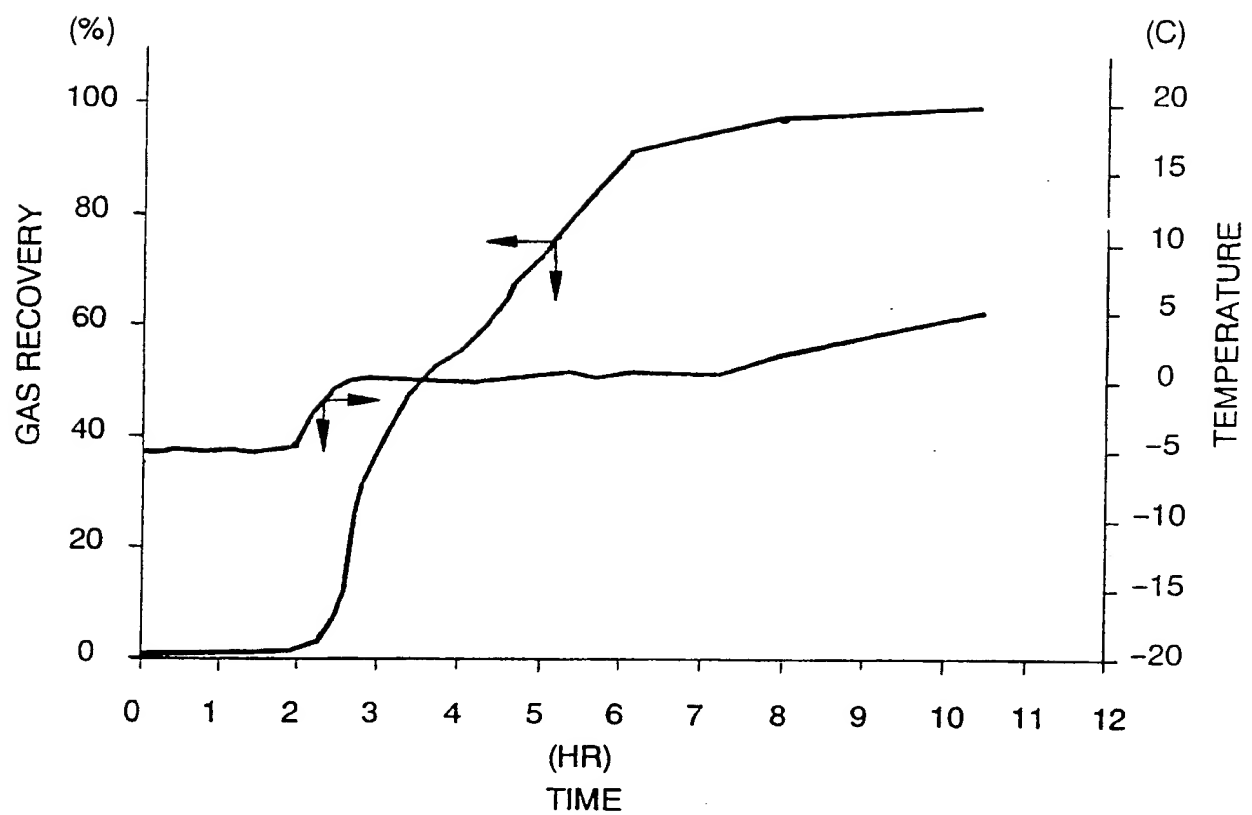


Fig. 6
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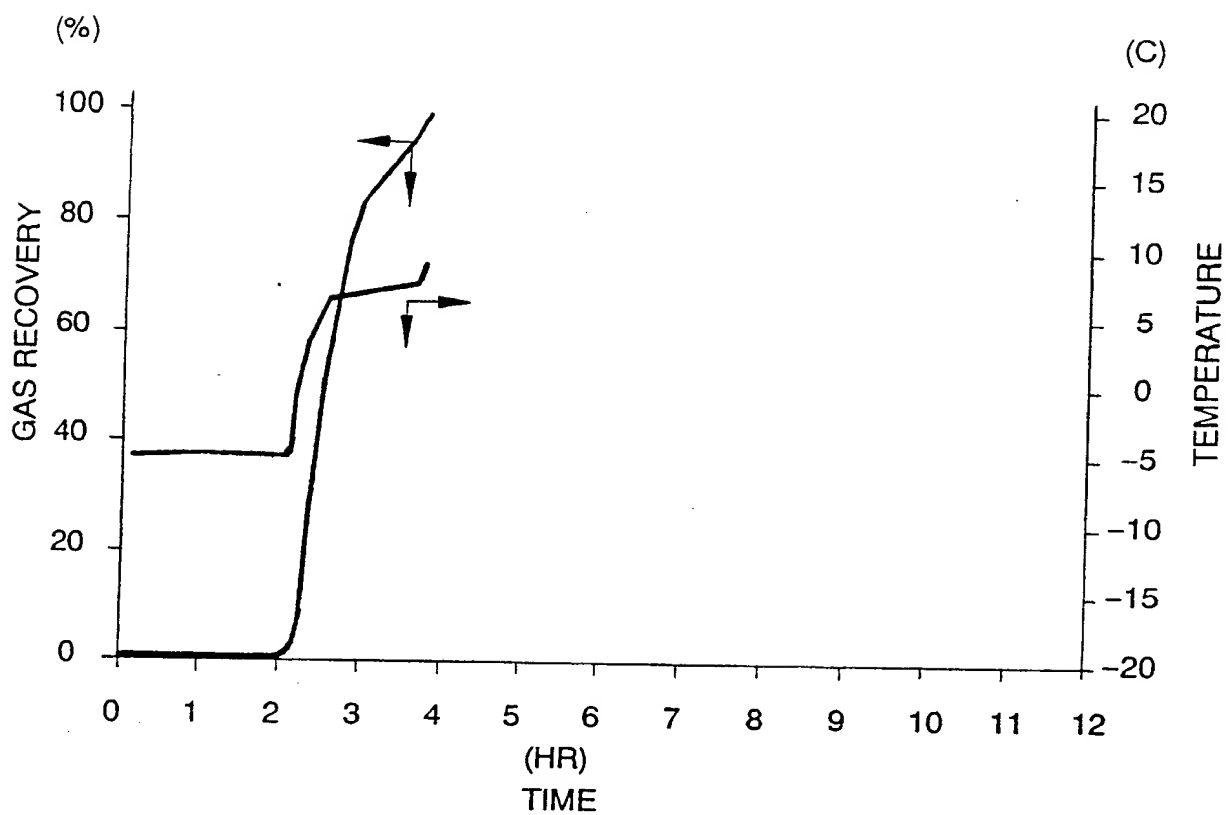
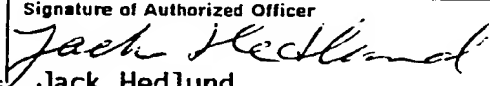
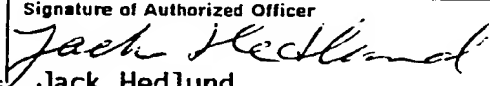
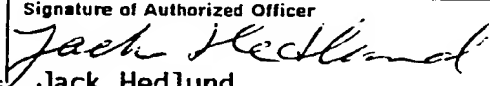


Fig. 7
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INTERNATIONAL SEARCH REPORT

International Application No PCT/NO 91/00101

I. CLASSIFICATION OF SUBJECT MATTER (if several classification symbols apply, indicate all) ⁶ According to International Patent Classification (IPC) or to both National Classification and IPC IPC5: C 07 C 7/20, F 25 J 1/02																	
II. FIELDS SEARCHED <div style="text-align: center; border-top: 1px solid black; border-bottom: 1px solid black;">Minimum Documentation Searched⁷</div> <table style="width: 100%; border-collapse: collapse;"> <tr> <th style="width: 20%; border: 1px solid black; text-align: left;">Classification System</th> <th style="border: 1px solid black; text-align: left;">Classification Symbols</th> </tr> <tr> <td style="border: 1px solid black; padding: 5px;">IPC5</td> <td style="border: 1px solid black; padding: 5px;">C 07 C; F 25 J</td> </tr> </table> <div style="text-align: center; border-top: 1px solid black; border-bottom: 1px solid black;">Documentation Searched other than Minimum Documentation to the Extent that such Documents are Included in Fields Searched⁸</div> <p style="padding: 5px;">SE,DK,FI,NO classes as above</p>			Classification System	Classification Symbols	IPC5	C 07 C; F 25 J											
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IPC5	C 07 C; F 25 J																
III. DOCUMENTS CONSIDERED TO BE RELEVANT⁹ <table style="width: 100%; border-collapse: collapse;"> <tr> <th style="width: 10%; border: 1px solid black; text-align: left;">Category *</th> <th style="border: 1px solid black; text-align: left;">Citation of Document,¹¹ with indication, where appropriate, of the relevant passages¹²</th> <th style="width: 15%; border: 1px solid black; text-align: left;">Relevant to Claim No.¹³</th> </tr> <tr> <td style="border: 1px solid black; text-align: center; vertical-align: top;">A</td> <td style="border: 1px solid black; padding: 5px;">US, A, 3514274 (R.P. CAHN ET AL) 26 May 1970, see the whole document --</td> <td style="border: 1px solid black; text-align: center; vertical-align: top;">1-11</td> </tr> <tr> <td style="border: 1px solid black; text-align: center; vertical-align: top;">A</td> <td style="border: 1px solid black; padding: 5px;">US, A, 3888434 (CHERSKY ET AL) 10 June 1975, see the whole document --</td> <td style="border: 1px solid black; text-align: center; vertical-align: top;">1-11</td> </tr> <tr> <td style="border: 1px solid black; text-align: center; vertical-align: top;">A</td> <td style="border: 1px solid black; padding: 5px;">US, A, 3975167 (NIERMAN) 17 August 1976, see the whole document --</td> <td style="border: 1px solid black; text-align: center; vertical-align: top;">1-11</td> </tr> <tr> <td style="border: 1px solid black; text-align: center; vertical-align: top;">A</td> <td style="border: 1px solid black; padding: 5px;">US, A, 4920752 (EHRSAM) 1 May 1990, see the whole document -- -----</td> <td style="border: 1px solid black; text-align: center; vertical-align: top;">1-11</td> </tr> </table>			Category *	Citation of Document, ¹¹ with indication, where appropriate, of the relevant passages ¹²	Relevant to Claim No. ¹³	A	US, A, 3514274 (R.P. CAHN ET AL) 26 May 1970, see the whole document --	1-11	A	US, A, 3888434 (CHERSKY ET AL) 10 June 1975, see the whole document --	1-11	A	US, A, 3975167 (NIERMAN) 17 August 1976, see the whole document --	1-11	A	US, A, 4920752 (EHRSAM) 1 May 1990, see the whole document -- -----	1-11
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A	US, A, 4920752 (EHRSAM) 1 May 1990, see the whole document -- -----	1-11															
<div style="display: flex; justify-content: space-between;"> <div style="width: 45%;"> <p>* Special categories of cited documents:¹⁰</p> <p>"A" document defining the general state of the art which is not considered to be of particular relevance</p> <p>"E" earlier document but published on or after the international filing date</p> <p>"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)</p> <p>"O" document referring to an oral disclosure, use, exhibition or other means</p> <p>"P" document published prior to the international filing date but later than the priority date claimed</p> </div> <div style="width: 45%;"> <p>"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention</p> <p>"X" document of particular relevance, the claimed invention cannot be considered novel or cannot be considered to involve an inventive step</p> <p>"Y" document of particular relevance, the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.</p> <p>"&" document member of the same patent family</p> </div> </div>																	
IV. CERTIFICATION <table style="width: 100%; border-collapse: collapse;"> <tr> <td style="width: 50%; border: 1px solid black; padding: 5px;"> Date of the Actual Completion of the International Search 11th February 1992 </td> <td style="width: 50%; border: 1px solid black; padding: 5px;"> Date of Mailing of this International Search Report 1992 -02- 12 </td> </tr> <tr> <td style="border: 1px solid black; padding: 5px;"> International Searching Authority <div style="text-align: center;">SWEDISH PATENT OFFICE</div> </td> <td style="border: 1px solid black; padding: 5px;"> Signature of Authorized Officer <div style="text-align: center;">  Jack Hedlund </div> </td> </tr> </table>			Date of the Actual Completion of the International Search 11th February 1992	Date of Mailing of this International Search Report 1992 -02- 12	International Searching Authority <div style="text-align: center;">SWEDISH PATENT OFFICE</div>	Signature of Authorized Officer <div style="text-align: center;">  Jack Hedlund </div>											
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ANNEX TO THE INTERNATIONAL SEARCH REPORT
ON INTERNATIONAL PATENT APPLICATION NO. PCT/NO 91/00101

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Patent document cited in search report	Publication date	Patent family member(s)	Publication date
US-A- 3514274	70-05-26	NONE	
US-A- 3888434	75-06-10	CA-A- 1018823 DE-A- 2411611	77-10-11 74-09-26
US-A- 3975167	76-08-17	AU-B- 500695 AU-D- 1257776 CA-A- 1034897 CA-A- 1056327 GB-A- 1534768 GB-A- 1534769	79-05-31 77-10-06 78-07-18 79-06-12 78-12-06 78-12-06
US-A- 4920752	90-05-01	CH-A- 677618 SE-A- 8900140	91-06-14 89-01-16

